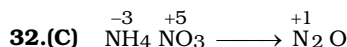
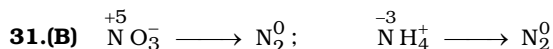
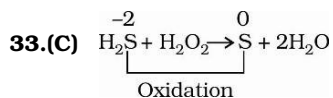


Daily Tutorial Sheet-3

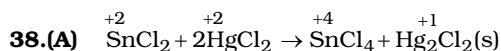
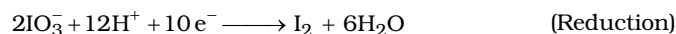
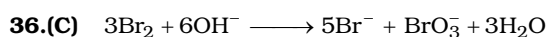
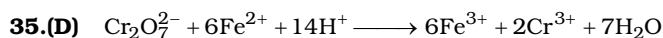
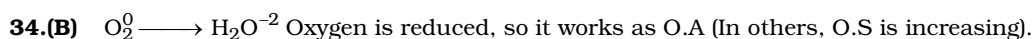
Level-1



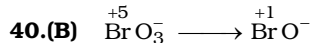
Since same atom is not reduced & oxidised.



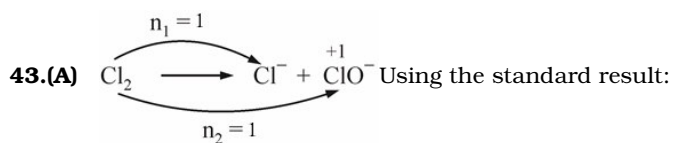
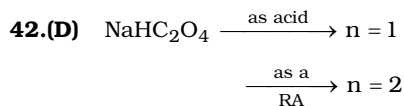
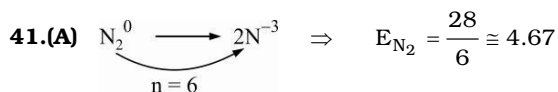
The oxidation of S shows oxidising nature of H_2O_2 . Also observe that oxidation state of O-atom in H_2O_2 increases from -1 to -2 in H_2O .



39.(B) SO_2 bleaches by reduction while chlorine bleaches colour of flowers by oxidation.

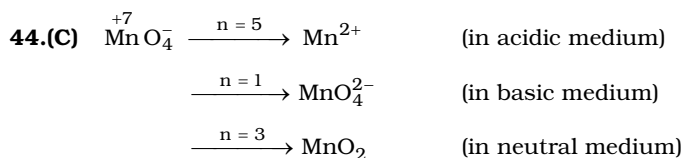


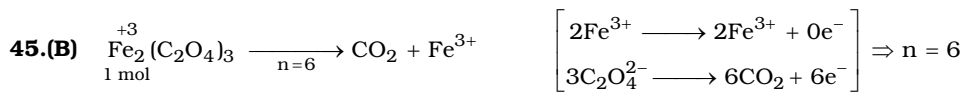
It is reduction, so needs a reducing agent.



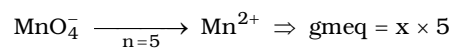
$$\Rightarrow \frac{1}{n} = \frac{1}{n_1} + \frac{1}{n_2} = \frac{1}{2} + \frac{1}{2} \Rightarrow n=1$$

Otherwise observe that 2 moles of $-e$ are involved per 2 mol of $\text{Cl}_2 \Rightarrow n=1$.





$$\Rightarrow \text{gmeq} = 6 \times 1$$



$$\text{g meq of Fe}_2(\text{C}_2\text{O}_4)_3 = \text{g meq of MnO}_4^-$$

$$1 \times 6 = x \times 5 \Rightarrow x = \frac{6}{5}$$